

Homework 3

$$\textcircled{1} \textcircled{i} u_k = \sqrt{g H S} = \sqrt{(9.8)(0.75)\left(\frac{1}{2500}\right)} = 0.0542 \text{ m s}^{-1}$$

$$D_y = 0.15 u_k H = (0.15)(0.0542)(0.75) = 6.1 \times 10^{-3} \text{ m}^2 \text{ s}^{-1}$$

$$D_z = 0.067 u_k H = (0.067)(0.0542)(0.75) = 2.724 \times 10^{-3} \text{ m}^2 \text{ s}^{-1}$$

$$\tau_y = \frac{w^2}{D_y} = \frac{q^2}{6.1 \times 10^{-3}} = 2624 \text{ s} \quad \tau_z = \frac{H^2}{D_z} = \frac{(0.75)^2}{2.72 \times 10^{-3}} = 206.5 \text{ s}$$

$$\tau_y > \tau_z = \tau = 2624 \text{ s}$$

$$u_{avg} = \frac{u_k \left(H \ln\left(\frac{H}{h_0}\right) - H + h_0 \right)}{K H} = \frac{(0.05) \left(0.75 \left(\ln\left(\frac{0.75}{0.015}\right) \right) - 0.75 + 0.015 \right)}{0.9 (0.75)} = 0.4 \text{ m/s}$$

$$l_{crit} = \tau u_{avg} = 2624 (0.4) = 1050 \text{ m}$$

$$450 \text{ m} < l_{crit} \Rightarrow \boxed{\text{Not well mixed}}$$

$$\textcircled{ii} C = \frac{M}{(4\pi D_x t)^{1/2}} e^{-\frac{(x - u_{avg} t)^2}{4 D_x t}}$$

$$D_x = \frac{0.011 u_{avg}^2 w^2}{u_k H} = 0.68 \text{ m}^2/\text{s}$$

$$\rightarrow M = C (4\pi D_x t)^{1/2} e^{+\frac{(x - u t)^2}{4 D_x t}}$$

$$t = \frac{5000}{u_{avg}}$$

$$M = \frac{(100 \times 10^{-6}) \left(4\pi (0.68) \left(\frac{5000}{0.4} \right) \right)^{1/2}}{e^{-\frac{(5000-5000)}{4D_x t}}} = \boxed{0.033 \text{ g}}$$

$$\sigma_{width} = \sqrt{2D_x t} = \sqrt{2(0.68) \left(\frac{5000}{0.4} \right)} = \boxed{130 \text{ m}}$$

② (i) $\gamma = 0.2 \text{ hr}^{-1} = 5.6 \times 10^{-5} \text{ s}^{-1}$

$$C = \frac{M}{(4\pi D_x t)^{1/2}} e^{-\frac{(x-ut)^2}{4D_x t}} e^{-\gamma t}$$

$$C_{peak} = \frac{M}{(4\pi D_x t)^{1/2}} e^{-\gamma t}$$

$$t = \frac{5000}{0.4} = 12500 \text{ s}$$

By design from ①

$$= 100 \mu\text{g} e^{-(5.6 \times 10^{-5} \times 12500)}$$

$$= 50 \mu\text{g/m}$$

(ii) Steady State

$$C = C_0 e^{\frac{u - \sqrt{u^2 + 4\gamma D_z}}{2D} x}$$

$$u = 0.4 \text{ ms}^{-1}$$

$$\gamma = 5.6 \times 10^{-5} \text{ s}^{-1}$$

$$D_z = 0.682 \text{ m}^2 \text{ s}^{-1}$$

$$x = 5000 \text{ m}$$

$$C_0 = 25 \text{ g/l}$$

$$\therefore C = 12.4 \text{ g/l} \leadsto \text{consistent with (i)} \leadsto \text{i.e. about } \underline{h_0/A}$$

$$(3) \text{ (i) } R = \int_0^H r dz = \int_0^H \alpha e^{-\lambda z} dz = -\frac{\alpha}{\lambda} (e^{-\lambda H} - 1)$$

$$\therefore R = -\frac{0.1}{10} (e^{-(10)(2.5)} - 1) = 0.01 \text{ g min}^{-1}$$

$$= 14.4 \text{ g day}^{-1}$$

$$(ii) C = C_0 + \frac{\alpha}{D_z \lambda} \left(z + \frac{1}{\lambda} e^{-\lambda z} \right) - \frac{\alpha}{D_z \lambda} \left(H + \frac{1}{\lambda} e^{-\lambda H} \right)$$

$$@ z=0 \Rightarrow C = 10 \text{ g/m} + \frac{0.1}{D_z (10)} \left(0 + \frac{1}{10} e^0 \right) - \frac{0.1}{D_z (10)} \left(2.5 + \frac{1}{10} e^{-10(2.5)} \right)$$

$$\text{Assume } D_z = 0.1 \text{ m}^2/\text{s} \Rightarrow \boxed{C = 9.76 \text{ g/m}}$$

$\hookrightarrow$ Ansok to specify

(iii) If sunlight doubles $\Rightarrow$ reaction rate $\propto$ doubles

(iv) Zeroth order reaction o.k. because C varies very little in the water column

$$9.76 \leq C \leq 10 \text{ g/m}$$

(4)

$$\text{Layer 1} \quad V_1 \frac{dC_1}{dt} = \cancel{Q_{gw_1} C_{gw_1}^0} + Q_{sw_1} C_{sw_1} + \cancel{Q_{sw_2} C_{sw_2}^0} + Q_{12}(C_2 - C_1) - Q_R C_1 - Q_{gw} C_1$$

$$8000 \frac{dC_1}{dt} = 10 C_{s1} + 3(C_2 - C_1) - 20 C_1 - 2 C_1$$

$$\boxed{8000 \frac{dC_1}{dt} = 10 C_{s1} - 25 C_1 + 3 C_2} \quad (A)$$

$$\text{Layer 2} \quad V_2 \frac{dC_2}{dt} = \cancel{Q_{gw_2} C_{gw_2}^0} + Q_{12}(C_1 - C_2) + Q_{23}(C_3 - C_2) - Q_{gw} C_2$$

$$7000 \frac{dC_2}{dt} = 3(C_1 - C_2) + 0.5(C_3 - C_2) - 2 C_2$$

$$\boxed{7000 \frac{dC_2}{dt} = 3 C_1 - 5.5 C_2 + 0.5 C_3} \quad (B)$$

Layer 3

$$V_3 \frac{dC_3}{dt} = Q_{cw3} \nearrow_{cw_{in}} + Q_{23} (C_2 - C_3) - Q_{cw} C_3$$

$$5000 \frac{dC_3}{dt} = 0.5 (C_2 - C_3) - 2C_3$$

$$\boxed{5000 \frac{dC_3}{dt} = 0.5 C_2 - 2.5 C_3} \quad (C)$$

$$\text{Steady state} \Rightarrow \frac{dC_1}{dt} = \frac{dC_2}{dt} = \frac{dC_3}{dt} = 0$$

$$\therefore \text{From (C)} \Rightarrow 0 = 0.5 C_2 - 2.5 C_3 \Rightarrow \boxed{C_2 = 5 C_3}$$

$$(B) \Rightarrow 3C_1 - 5.5C_2 + 0.5C_3 = 0$$

$$3C_1 = 5.5C_2 - 0.1C_2 = 5.6C_2$$

$$\boxed{C_1 = \frac{5.6}{3} C_2}$$

$$(A) \quad 10 C_{\text{stream}} - 25C_1 + 3C_2 = 0$$

$$\boxed{C_{\text{stream}} = \frac{25C_1 - 3C_2}{10}}$$

$$C_3 = 1 \text{ g/l} \Rightarrow C_2 = 5 \text{ g/l} \Rightarrow C_1 = 9 \text{ g/l}$$

$$\Rightarrow \boxed{C_{\text{stream}} = 21 \text{ g/l}}$$

Final part numerical !